18. What do atoms do to complete their outer level of electrons? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

19. Positive oxidation numbers mean that the atom did what in order to complete the octet rule?

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

20. Opposite charges \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ each other while like charges \_\_\_\_\_\_\_\_\_\_\_\_\_\_ each other.

21. If an ion has a negative oxidation number, more than likely what kind of element was this?

 Metal or Nonmetal (circle one)

22. When atoms bond covalently, the resulting particle is called a/an \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

23. This number tells how many electrons an atom gains, loses or shares in forming compounds:

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

24. In writing formulas, the sum of all the oxidation number for the atoms in the compounds has to

 equal \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

25. In writing ionic formulas, which one comes first? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Name the following compounds from the chemical formulas

1. Ca3P2
2. C2O6
3. MgS
4. N2O5
5. SO2
6. AlCl3

Write the formulas for each of the following compounds

1. Calcium Bromide
2. Sodium Nitride
3. Dinitrogen Tetroxide
4. Sulfur Trioxide
5. Triphosphorous Heptachloride
6. Boron Nitride